

Name: \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_ Score: \_\_\_\_\_/3

### SUPERMARKET CHEMISTRY

NAME AND FORMULA GRID – Applying polyatomic ions in ionic compound formula writing and naming

	Oxide O <sup>2-</sup>	Chloride Cl <sup>-</sup>	Phosphate	Sulfide	Sulfate	Hydroxide	Carbonate CO <sub>3</sub> <sup>2-</sup>	Acetate
H <sup>+</sup> <b>hydrogen ion</b>	H <sub>2</sub> O water	HCl hydrogen chloride						
NH <sub>4</sub> <sup>+</sup> <b>ammonium ion</b>		<b>NH<sub>4</sub>Cl</b> <b>ammonium</b> <b>chloride</b>						
Na <sup>+</sup>							<b>Na<sub>2</sub>CO<sub>3</sub></b> <b>sodium</b> <b>carbonate</b>	
Ca <sup>+2</sup>								<b>Ca(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub></b> <b>calcium</b> <b>acetate</b>
Fe <sup>+2</sup>				<b>FeS</b> <b>iron(II) sulfide</b>				
Fe <sup>+3</sup>						<b>Fe(OH)<sub>3</sub></b> <b>iron(III)</b> <b>hydroxide</b>		
Al <sup>+3</sup>					<b>Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b> <b>aluminum</b> <b>sulfate</b>			
Sn <sup>+4</sup> <b>tin (IV)</b>		<b>SnCl<sub>4</sub></b> <b>tin(IV)</b> <b>chloride</b>						

Write the name and formula for each of the ions (including Roman Numerals where appropriate) in each of the boxes provided. In each of the boxes where the anions and cations meet, write the neutral balanced formula followed by the correct compound name.