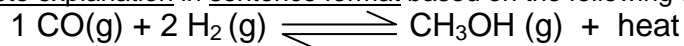


Le Chatelier's Principle: When a stress is applied to a system at equilibrium the system will adjust to relieve the stress.

Based on your understanding of Le Chatelier's Principle, predict the response to the following changes to restore equilibrium by identifying the shift (right or left) to production of more products or more reactants. Use diagrams to accompany your full and complete explanation in sentence format based on the following chemical equilibrium reaction:



1. CO (g) is added to the system:

2. H₂ gas is removed:

3. Temperature increases when the reaction is heated:

4. Volume is increased:

5. Pressure is increased:

BENCHMARK ACTIVITY: CA Content Standard: **Chemical Equilibrium**

9. Chemical equilibrium is a dynamic process at the molecular level. As a basis for understanding this concept, students know:
- how to use LeChatelier's Principle to predict the effect of changes in concentration, temperature and pressure.
 - equilibrium is established when forward and reverse reaction rates are equal.