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Sweet 16 Chemistry Compound Tournament

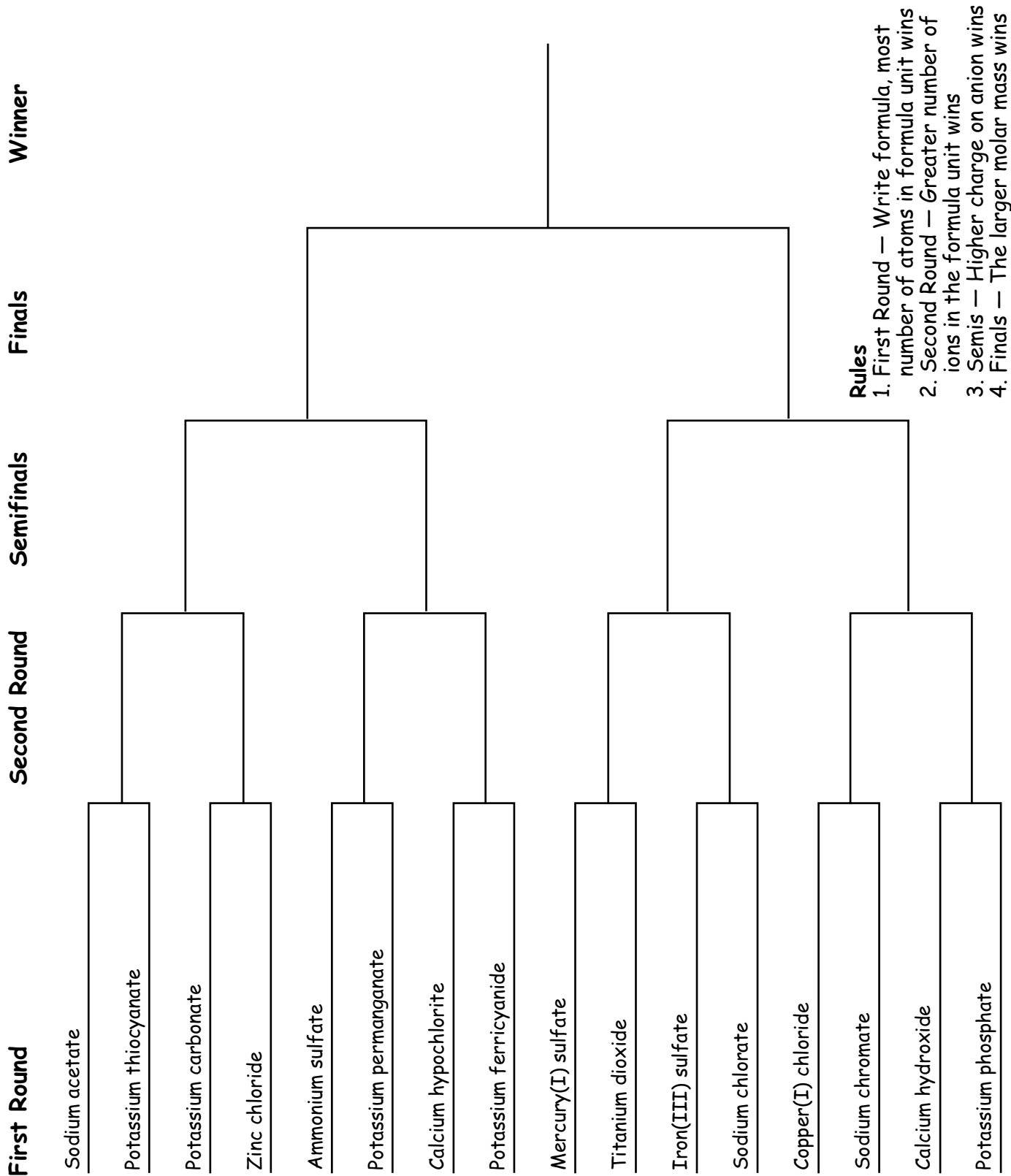
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With spring just around the corner, your students' thoughts will soon be turning to sunshine, prom, and the NCAA basketball tournament. This clever activity combines the ever-popular March Madness basketball pool with a review of chemical formulas, the structure and charges of ions, and molar mass calculations. The result is the Sweet 16 Chemistry Compound Tournament. Hopefully, your students will enjoy playing the "Tournament" while reviewing these important topics.

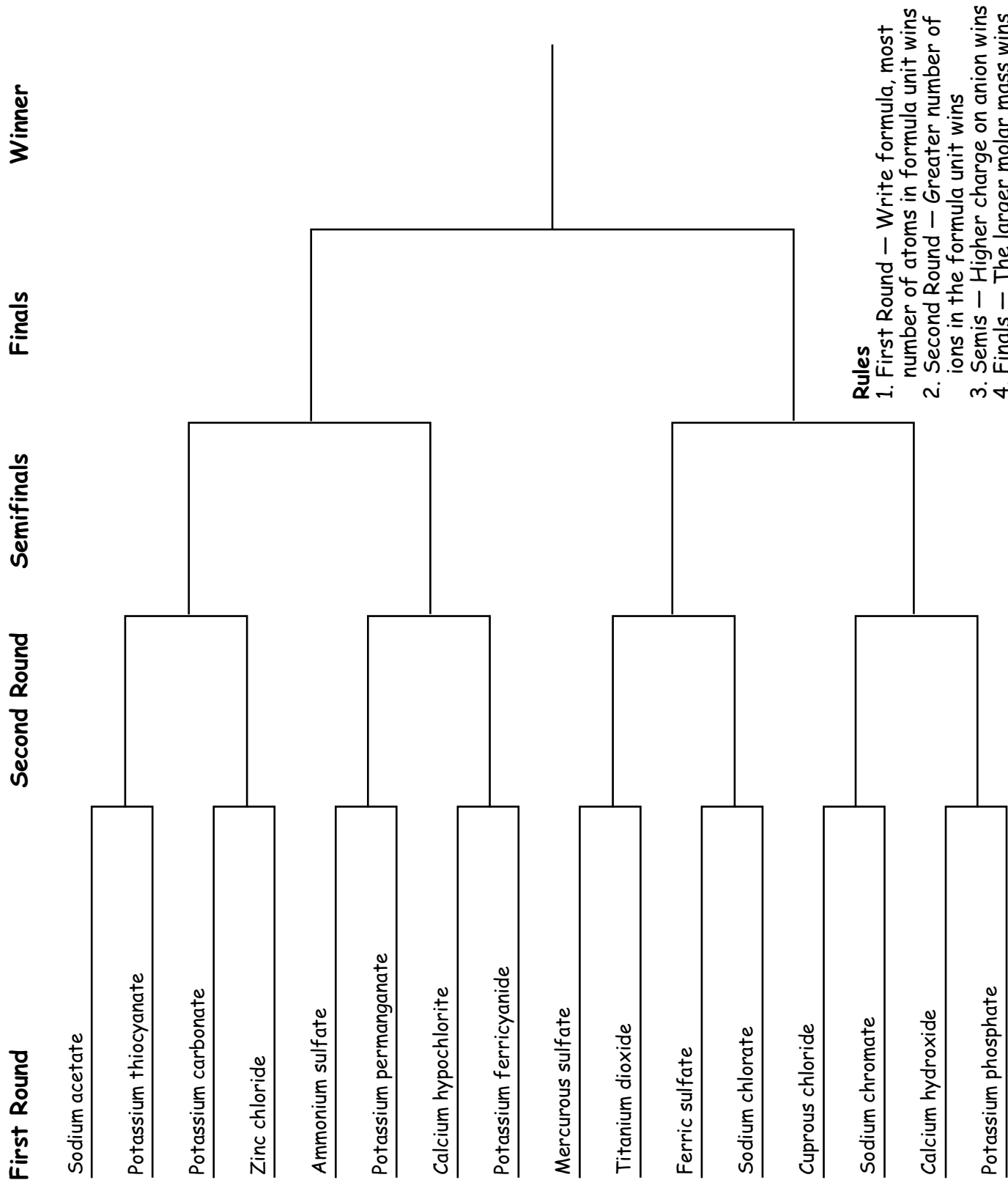
The rules for filling out the Tournament brackets are simple: For the first round, the student determines the chemical formula for the compound and counts the total number of atoms in the formula unit. The compound with the greater number of atoms in the formula unit wins and moves on to the second round. In the second round, the two compounds in each contest are placed in water and the compound that produces the greater number of ions wins this round. In the semifinals, the compound containing the anion with the higher charge wins the contest and moves on to the finals. The winner of the tournament is the compound in the finals that has the larger molar mass.

Two versions of the Sweet 16 Chemistry Compound Tournament are enclosed. One uses traditional chemical names [e.g. ferrous, cupric, etc] and the other uses the Stock system for chemical names [iron(II), copper(II)]. If you have already received your 2005 *Flinn Scientific Catalog/Reference Manual*, you may have noticed that we have updated all our chemical names, labels, and MSDS to the more familiar Stock system of nomenclature. This change was a result of many chemistry teachers asking us to change to the Stock system. If you have not received a 2005 *Flinn Scientific Catalog/Reference Manual*, please give us a call (800-452-1261) or e-mail (flinn@flinnsci.com) to request one.

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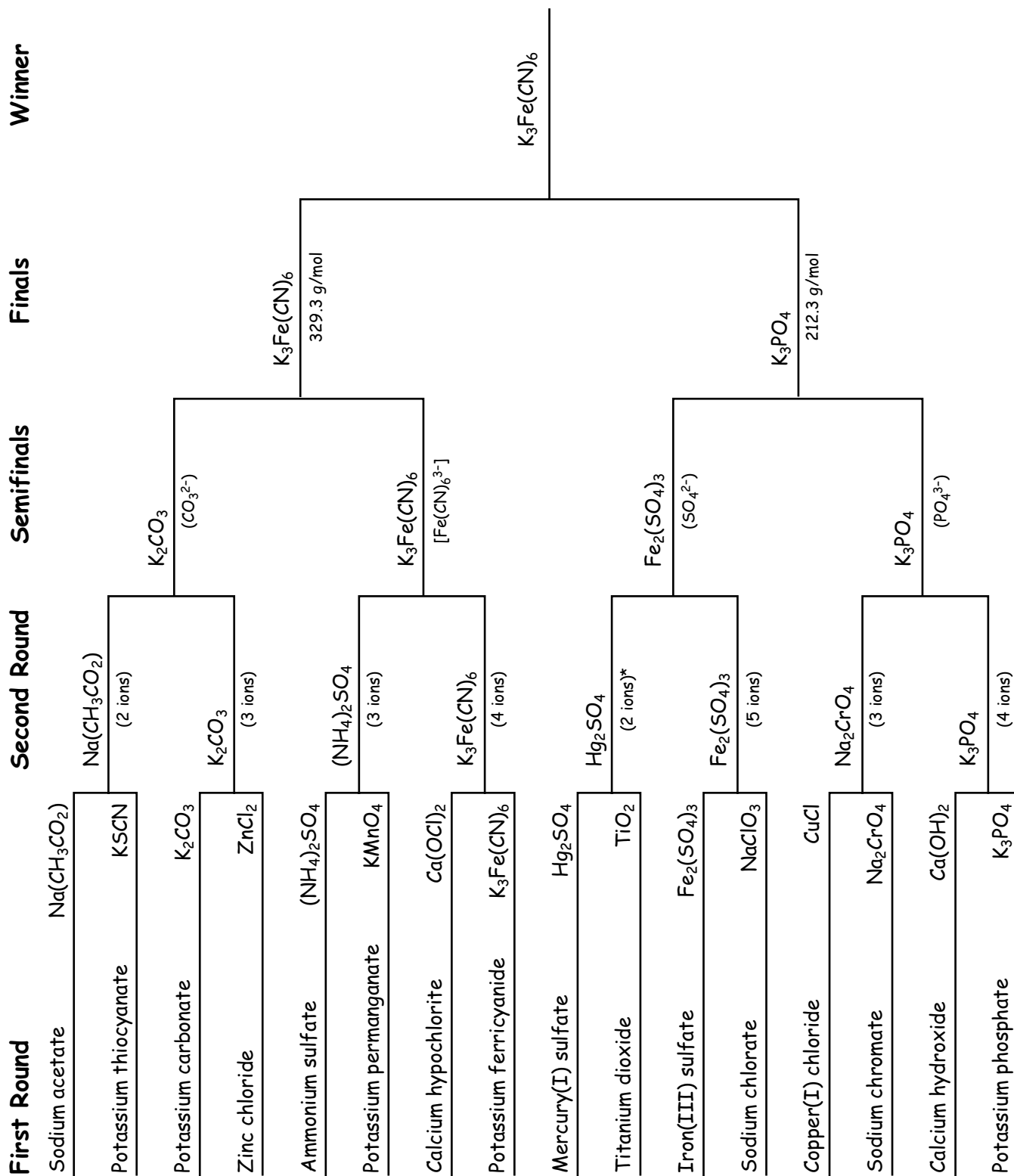


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Answer Key



*The mercury(I) ion exists as Hg₂²⁺.